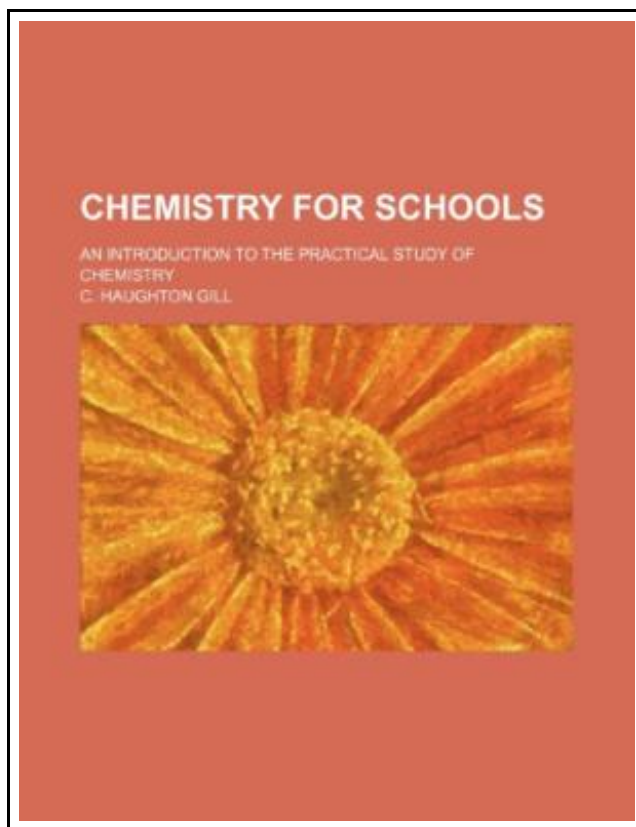


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1869 Excerpt: .of a very simple method, and apparatus for the purpose, will be found in Golding Bird s Elements of Natural Philosophy, p. 372. Ammonium amalgam may be prepared by another and simpler process, which the student can readily carry out. Prepare about four ounces of a saturated solution of ammoniac chloride in a basin. Then make a little sodium amalgam by throwing a bit of sodium the size of a large pea on to the surface of a quarter of an ounce of wet mercury in a test-tube (the two metals unite with a slight flash). Immediately pour the sodium amalgam into the ammoniac chloride. A rapid action at once commences; the mercury swells up to one or two hundred times its former bulk, but retains in a great measure its lustre. The swollen mass when removed from the liquid and exposed to the air or thrown into water decomposes in a short time with evolution of hydrogen and ammonia. The action, which doubtless takes place, is an interchange of sodium and ammonium, the first quitting the mercury to combine with chlorine, which the ammonium leaves for the mercury. $NH_4Cl + NaHg = NH_4Hg + NaCl$. Hgn means that it is an unknown multiple of the atomic weight of mercury which takes part in this reaction. 222, When ammonia gas is passed into water, it is supposed to form ammoniac hydrate, $NH_4(OH)$, corresponding to potassic hydrate, $K(OH)$,...



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